

1.	Mg reacts with HCl as per the following reaction: $\text{Mg}_{(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{MgCl}_{2(aq)} + \text{H}_{2(g)}$ Given that Mg = 24 g and HCl = 36.5 g, the excess reactant is: <input checked="" type="radio"/> (a) Mg (b) HCl (c) Both are in stoichiometric amounts (d) None of these
2.	When 0.01 kg of CaCO_3 is decomposed the CO_2 produced occupies a volume at S.T.P.: <input checked="" type="radio"/> (a) 2.2414 dm ³ (b) 22.414 dm ³ (c) 22414 dm ³ (d) 224014 dm ³
3.	A compound contain 50% sulphur and 50% oxygen by mass. The empirical formula of the compound is: <input checked="" type="radio"/> (a) S_2O_3 <input checked="" type="radio"/> (b) SO_2 (c) SO_3 (d) SO
4.	Combustion analysis is performed for the determination of: (a) Molar mass of the compound <input checked="" type="radio"/> (b) Empirical formula of the compound (c) Structural formula of the compound (d) Mass of halogen present in organic compound
5.	During combustion analysis, which one is used for absorbing carbon dioxide: <input checked="" type="radio"/> (a) 50% KOH solution (b) 5% KOH (c) $\text{Mg}(\text{ClO}_4)_2$ (d) Silica gel
6.	5604 cm ³ of H_2 gas at STP contains atoms of hydrogen: (a) 6.02×10^{23} (b) 2.6×10^{23} <input checked="" type="radio"/> (c) 3.01×10^{23} (d) 1.50×10^{23}
7.	% of nitrogen in urea (NH_2CONH_2): (a) 82.35% <input checked="" type="radio"/> (b) 6.6% (c) 35% (d) 55.56%
8.	Which of the following has highest percentage of nitrogen: (a) $(\text{NH}_4)_2\text{SO}_4$ (b) $\text{NH}_4\text{H}_2\text{PO}_4$ (c) $(\text{NH}_4)_2\text{HPO}_4$ <input checked="" type="radio"/> (d) $(\text{NH}_4)_3\text{PO}_4$
9.	500 cm ³ of H_2 gas at STP contain hydrogen: (a) 6.02×10^{23} (b) 3.01×10^{22} (c) 2.68×10^{22} <input checked="" type="radio"/> (d) 0.154×10^{22}
10.	Empirical formula of chloroform is (a) CH_2Cl_2 (b) CHCl_3 (c) CCl_4 <input checked="" type="radio"/> (d) CHCl_3
11.	The volume occupied by 2.8 g of N_2 at STP <input checked="" type="radio"/> (a) 2.24 dm ³ (b) 22.4 dm ³ (c) 1.12 dm ³ (d) 112 dm ³
12.	The mass of one molecule of O_2 is: (a) $6.02 \times 10^{23} / 32$ (b) 0.32 g (c) 32 g <input checked="" type="radio"/> (d) $32 / 6.02 \times 10^{23}$
13.	6.02×10^{23} atoms of tetra atomic gas at S.T.P occupy a volume of: (a) 2.9 Liter (b) 5.8 Liter (c) 224 Liter <input checked="" type="radio"/> (d) 22.414 Liter
14.	The mass of one mole of CO_3^{2-} is: (a) 40g (b) 600g <input checked="" type="radio"/> (c) 60g (d) 62g
15.	10g of Mg and 5g of C contain the same: (a) No. of moles (b) No. of atoms <input checked="" type="radio"/> (c) Both a & b (d) None
16.	The simplest formula of a substance is CH_2 and its molecular weight is 56, what is its actual formula: (a) CH_2 (b) C_2H_2 (c) C_7H_8 <input checked="" type="radio"/> (d) C_8H_8
17.	Mass of 0.4 moles ozone in comparison to the mass of 0.4 moles of O_2 are: (a) Same (b) Less <input checked="" type="radio"/> (c) Greater (d) None of these
18.	Which one of the following compound does not have the empirical formula CH_2O : (a) HCHO (b) $\text{C}_6\text{H}_{12}\text{O}_6$ (c) CH_3COOH <input checked="" type="radio"/> (d) $\text{CH}_3\text{CH}_2\text{OH}$

Mcat The Answer Key

Nilanjan Sen



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